SELECTED STANDARD REDUCTION POTENTIALS AT 298 K

Reduction half-reactions			**E°(V)**
$Li^{1+}(aq) + e^{1-}$	$\rightleftharpoons$	$Li(s)$	−3.04
$K^{1+}(aq) + e^{1-}$	$\rightleftharpoons$	$K(s)$	−2.92
$Ba^{2+}(aq) + 2\ e^{1-}$	$\rightleftharpoons$	$Ba(s)$	−2.92
$Ca^{2+}(aq) + 2\ e^{1-}$	$\rightleftharpoons$	$Ca(s)$	−2.84
$Na^{1+}(aq) + e^{1-}$	$\rightleftharpoons$	$Na(s)$	−2.71
$Mg^{2+}(aq) + 2\ e^{1-}$	$\rightleftharpoons$	$Mg(s)$	−2.36
$Al^{3+}(aq) + 3\ e^{1-}$	$\rightleftharpoons$	$Al(s)$	−1.66
$U^{3+}(aq) + 3\ e^{1-}$	$\rightleftharpoons$	$U(s)$	−1.66
$Ti^{2+}(aq) + 2\ e^{1-}$	$\rightleftharpoons$	$Ti(s)$	−1.63
$Mn^{2+}(aq) + 2\ e^{1-}$	$\rightleftharpoons$	$Mn(s)$	−1.18
$2\ H_2O + 2\ e^{1-}$	$\rightleftharpoons$	$H_2(g) + 2\ OH^{1-}(aq)$	−0.83
$Zn^{2+}(aq) + 2\ e^{1-}$	$\rightleftharpoons$	$Zn(s)$	−0.76
$Cr^{3+}(aq) + 3\ e^{1-}$	$\rightleftharpoons$	$Cr(s)$	−0.74
$HCHO(aq) + 2\ H_2O + 2\ e^{1-}$	$\rightleftharpoons$	$CH_3OH(aq) + 2\ OH^{1-}(aq)$	−0.59
$Fe^{2+}(aq) + 2\ e^{1-}$	$\rightleftharpoons$	$Fe(s)$	−0.44
$2\ H_2O + 2\ e^{1-}$	$\rightleftharpoons$	$H_2(g) + 2\ OH^{1-}(aq)$	−0.41*
$Cd^{2+}(aq) + 2\ e^{1-}$	$\rightleftharpoons$	$Cd(s)$	−0.40
$PbSO_4(s) + 2\ e^{1-}$	$\rightleftharpoons$	$Pb(s) + SO_4^{2-}(aq)$	−0.36
$In^{3+}(aq) + 3\ e^{1-}$	$\rightleftharpoons$	$In(s)$	−0.34
$Co^{2+}(aq) + 2\ e^{1-}$	$\rightleftharpoons$	$Co(s)$	−0.28
$Ni^{2+}(aq) + 2\ e^{1-}$	$\rightleftharpoons$	$Ni(s)$	−0.23
$Sn^{2+}(aq) + 2\ e^{1-}$	$\rightleftharpoons$	$Sn(s)$	−0.14
$Pb^{2+}(aq) + 2\ e^{1-}$	$\rightleftharpoons$	$Pb(s)$	−0.13
$2\ H^{1+}(aq) + 2\ e^{1-}$	$\rightleftharpoons$	$H_2(g)$	0.00
$Sn^{4+}(aq) + 2\ e^{1-}$	$\rightleftharpoons$	$Sn^{2+}(aq)$	+0.15
$Cu^{2+}(aq) + e^{1-}$	$\rightleftharpoons$	$Cu^{1+}(aq)$	+0.16
$ClO_4^{1-}(aq) + H_2O(l) + 2\ e^{1-}$	$\rightleftharpoons$	$ClO_3^{1-}(aq) + 2\ OH^{1-}(aq)$	+0.17
$AgCl(s) + e^{1-}$	$\rightleftharpoons$	$Ag(s) + Cl^{1-}(aq)$	+0.22
$PbO_2(s) + 2\ H^{1+}(aq) + 2\ e^{1-}$	$\rightleftharpoons$	$PbO(s) + H_2O$	+0.28
$Cu^{2+}(aq) + 2\ e^{1-}$	$\rightleftharpoons$	$Cu(s)$	+0.34
$Ag_2O(s) + H_2O + 2\ e^{1-}$	$\rightleftharpoons$	$2\ Ag(s) + 2\ OH^{1-}(aq)$	+0.34
$ClO_3^{1-}(aq) + H_2O(l) + 2\ e^{1-}$	$\rightleftharpoons$	$ClO_2^{1-}(aq) + 2\ OH^{1-}(aq)$	+0.35
$O_2(g) + 2\ H_2O + 4\ e^{1-}$	$\rightleftharpoons$	$4\ OH^{1-}(aq)$	+0.40
$I_2(s) + 2\ e^{1-}$	$\rightleftharpoons$	$2\ I^{1-}(aq)$	+0.54
$ClO_2^{1-}(aq) + H_2O(l) + 2\ e^{1-}$	$\rightleftharpoons$	$ClO^{1-}(aq) + 2\ OH^{1-}(aq)$	+0.59
$2\ AgO(s) + H_2O + 2\ e^{1-}$	$\rightleftharpoons$	$Ag_2O(s) + 2\ OH^{1-}(aq)$	+0.60
$O_2(g) + 2\ H^{1+}(aq) + 2\ e^{1-}$	$\rightleftharpoons$	$H_2O_2(aq)$	+0.70
$Fe^{3+}(aq) + e^{1-}$	$\rightleftharpoons$	$Fe^{2+}(aq)$	+0.77
$BrO^{1-}(aq) + H_2O + 2\ e^{1-}$	$\rightleftharpoons$	$Br^{1-}(aq) + 2\ OH^{1-}(aq)$	+0.77

Reduction half-reactions			**E°(V)**
$Ag^{1+}(aq) + e^{1-}$	$\rightleftharpoons$	$Ag(s)$	+0.80
$O_2(g) + 4\ H^{1+}(aq) + 4\ e^{1-}$	$\rightleftharpoons$	$2\ H_2O$	+0.82*
$H_2O_2(aq) + 2\ e^{1-}$	$\rightleftharpoons$	$2\ OH^{1-}(aq)$	+0.88
$ClO^{1-}(aq) + H_2O + 2\ e^{1-}$	$\rightleftharpoons$	$Cl^{1-}(aq) + 2\ OH^{1-}(aq)$	+0.89
$Hg^{2+}(aq) + 2\ e^{1-}$	$\rightleftharpoons$	$Hg(l)$	+0.85
$NO_3^{1-}(aq) + 4\ H^{1+}(aq) + 3\ e^{1-}$	$\rightleftharpoons$	$NO(g) + 2\ H_2O$	+0.96
$VO_2^{1+}(aq) + 2\ H^{1+}(aq) + e^{1-}$	$\rightleftharpoons$	$VO^{2+}(aq) + H_2O$	+1.00
$Br_2(l) + 2\ e^{1-}$	$\rightleftharpoons$	$2\ Br^{1-}(aq)$	+1.09
$ClO_4^{1-}(aq) + 2\ H^{1+}(aq) + 2\ e^{1-}$	$\rightleftharpoons$	$ClO_3^{1-}(aq) + H_2O(l)$	+1.19
$O_2(g) + 4\ H^{1+}(aq) + 4\ e^{1-}$	$\rightleftharpoons$	$2\ H_2O$	+1.23
$Cr_2O_7^{2-}(aq) + 14\ H^{1+}(aq) + 6\ e^{1-}$	$\rightleftharpoons$	$2\ Cr^{3+}(aq) + 7\ H_2O$	+1.33
$Cl_2(g) + 2\ e^{1-}$	$\rightleftharpoons$	$2\ Cl^{1-}(aq)$	+1.36
$Au^{3+}(aq) + 3\ e^{1-}$	$\rightleftharpoons$	$Au(s)$	+1.50
$MnO_4^{1-}(aq) + 8\ H^{1+}(aq) + 5\ e^{1-}$	$\rightleftharpoons$	$Mn^{2+}(aq) + 4\ H_2O$	+1.51
$PbO_2(s) + 4\ H^{1+}(aq) + SO_4^{2-}(aq) + 2\ e^{1-}$	$\rightleftharpoons$	$PbSO_4(s) + 2\ H_2O$	+1.69
$H_2O_2(aq) + 2\ H^{1+}(aq) + 2\ e^{1-}$	$\rightleftharpoons$	$2\ H_2O(l)$	+1.76
$S_2O_8^{2-}(aq) + 2\ e^{1-}$	$\rightleftharpoons$	$2\ SO_4^{2-}(aq)$	+2.01
$O_3(g) + 2\ H^{1+}(aq) + 2\ e^{1-}$	$\rightleftharpoons$	$O_2(g) + H_2O$	+2.07
$F_2(g) + 2\ e^{1-}$	$\rightleftharpoons$	$2\ F^{1-}(aq)$	+2.87

* The half-cell potentials for the O_2/H_2O and the H_2O/H_2 systems are for neutral (pH = 7) water and are not standard reduction potentials where $[OH^{1-}]$ or $[H^{1+}] = 1.0$ M.