Metals That Form Cations with the Same Charge		
Metal Name	***Ion***	***Ion Name***
lithium	Li^+	lithium ion
sodium	Na^+	sodium ion
potassium	K^+	potassium ion
rubidium	Rb^+	rubidium ion
cesium	Cs^+	cesium ion
beryllium	Be^{2+}	beryllium ion
magnesium	Mg^{2+}	magnesium ion
calcium	Ca^{2+}	calcium ion
strontium	Sr^{2+}	strontium ion
barium	Ba^{2+}	barium ion
aluminum	Al^{3+}	aluminum ion
zinc	Zn^{2+}	zinc ion
scandium	Sc^{3+}	scandium ion
silver	Ag^+	silver ion

Metals That Form Cations with Different Charges			
Metal Name	***Ion***	***Ion Name***	***Historic Name***
chromium	Cr^{2+}	chromium(II) ion	chromous ion
	Cr^{3+}	chromium(III) ion	chromic ion
iron	Fe^{2+}	iron(II) ion	ferrous ion
	Fe^{3+}	iron(III) ion	ferric ion
cobalt	Co^{2+}	cobalt(II) ion	cobaltous ion
	Co^{3+}	cobalt(III) ion	cobaltic ion
copper	Cu^+	copper(I) ion	cuprous ion
	Cu^{2+}	copper(II) ion	cupric ion
tin	Sn^{2+}	tin(II) ion	stannous ion
	Sn^{4+}	tin(IV) ion	stannic ion
mercury	Hg_2^{2+}	mercury(I) ion	mercurous ion
	Hg^{2+}	mercury(II) ion	mercuric ion
lead	Pb^{2+}	lead(II) ion	plumbous ion
	Pb^{4+}	lead(IV) ion	plumbic ion

Monatomic Anions		
Non-Metal Name	***Ion***	***Ion Name***
fluorine	F^-	fluoride ion
chlorine	Cl^-	chloride ion
bromine	Br^-	bromide ion
iodine	I^-	iodide ion
oxygen	O^{2-}	oxide ion
sulfur	S^{2-}	sulfide ion
nitrogen	N^{3-}	nitride ion
phosphorous	P^{3-}	phosphide ion

Common Polyatomic Ions	
Ion	***Ion Name***
NH_4^+	ammonium ion
$C_2H_3O_2^-$ or $CH_3CO_2^-$	acetate ion
CN^-	cyanide ion
OH^-	hydroxide ion
NO_3^-	nitrate ion
NO_2^-	nitrite ion
ClO_4^-	perchlorate ion
ClO_3^-	chlorate ion
ClO_2^-	chlorite ion
ClO^-	hypochlorite ion
MnO_4^-	permanganate ion
HCO_3^-	hydrogen carbonate or bicarbonate ion
CO_3^{2-}	carbonate ion
O_2^{2-}	peroxide ion
HSO_4^-	hydrogen sulfate or bisulfate ion
SO_4^{2-}	sulfate ion
SO_3^{2-}	sulfite ion
$H_2PO_4^-$	dihydrogen phosphate ion
HPO_4^{2-}	hydrogen phosphate ion
PO_4^{3-}	phosphate ion